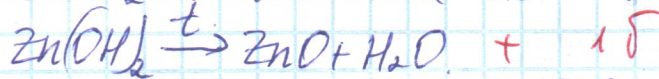
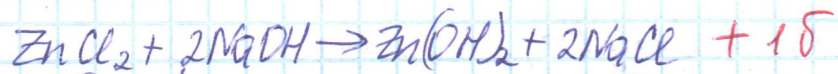


N 11.3

Нагрузка 10

36,50



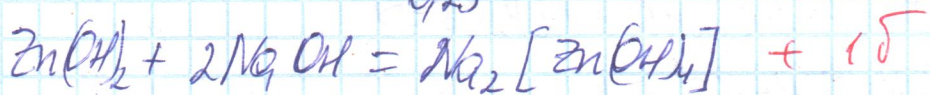
$$n(\text{ZnCl}_2) = \frac{m}{M} = \frac{275 \cdot 0,1088}{136} = 0,22 \text{ моль} +$$

$$n(\text{ZnO}) = \frac{m}{M} = \frac{8,1}{81} = 0,1 \text{ моль} + 15$$

$$n(\text{NaOH}) = 2 \cdot n(\text{Zn(OH)}_2) = 2 \cdot 0,1 = 0,2 \text{ моль}$$

$$n(\text{Zn(OH)}_2)_1 = n(\text{ZnO}) = 0,1 \text{ моль}$$

$$c(\text{NaOH}) = \frac{n}{V} = \frac{0,2}{0,25} = 0,8 \text{ моль/л} +$$



$$n(\text{Zn(OH)}_2)_2 = n(\text{ZnCl}_2) = 0,22 \text{ моль}$$

$$n(\text{Zn(OH)}_2)_3 = 0,22 - 0,1 = 0,12 \text{ моль} +$$

$$n(\text{NaOH})_1 = 2 \cdot n(\text{Zn(OH)}_2)_2 = 2 \cdot 0,22 = 0,44 \text{ моль}$$

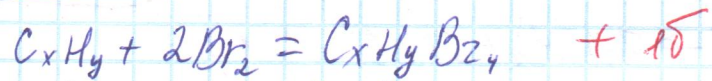
$$n(\text{NaOH})_3 = 2 \cdot n(\text{Zn(OH)}_2)_3 = 2 \cdot 0,12 = 0,24 \text{ моль}$$

$$n(\text{NaOH})_{\text{общ}} = 0,44 + 0,24 = 0,68 \text{ моль} +$$

$$c(\text{NaOH}) = \frac{n_{\text{общ}}}{V} = \frac{0,68}{0,25} = 2,72 \text{ моль/л} +$$

80

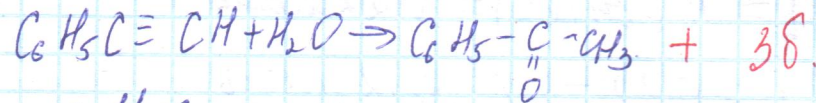
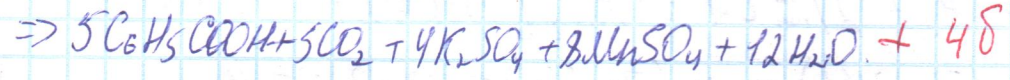
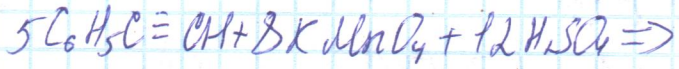
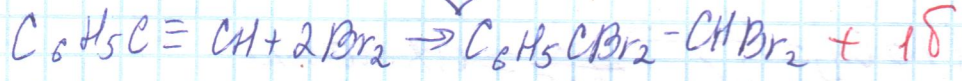
N 11-1



$$M(C_xH_yBr_4) = \frac{80 \cdot 4}{0,758} = 422 \text{ г/моль}$$

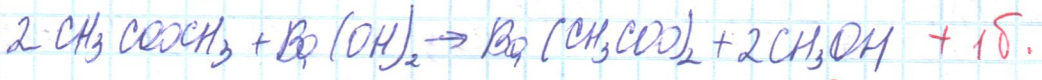
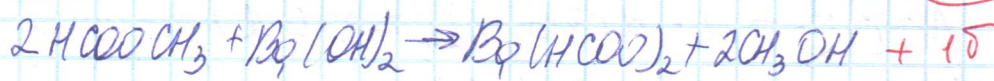
$$M(C_xH_y) = M(C_xH_yBr_4) - M(2Br_2) = 422 - (2 \cdot 160) = 102 \text{ г/моль} + 40.$$

Значит, формула  $C_8H_2$



150

N 11-2



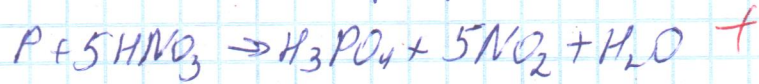
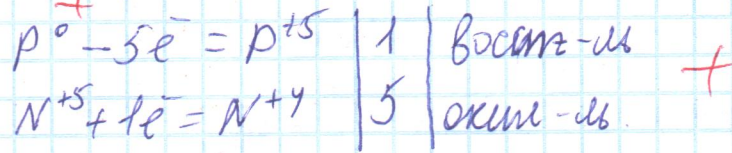
$$20,8 = 60x + 74y \quad 0,50.$$

3,50

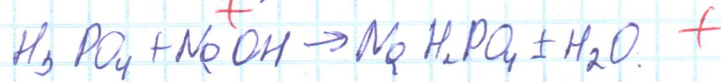


N11-5

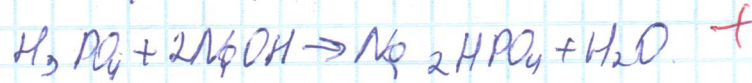
A-P, B-HNO<sub>3</sub> +



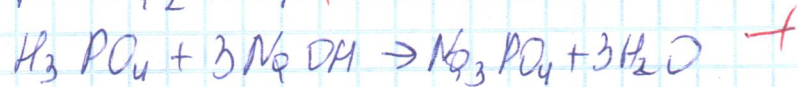
B - H<sub>3</sub>PO<sub>4</sub>, 2e - NO<sub>2</sub> +



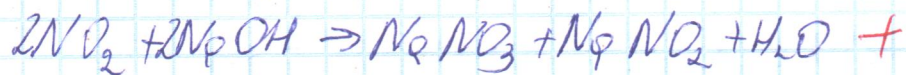
Г - NO<sub>2</sub>H<sub>2</sub>PO<sub>4</sub> +



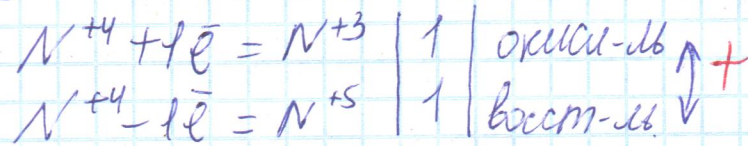
Д - NO<sub>2</sub>HPO<sub>4</sub> +



Е - NO<sub>2</sub>PO<sub>4</sub> +



Ж - NO<sub>2</sub>NO<sub>3</sub> +, К - NO<sub>2</sub>NO<sub>2</sub> +



100